

تم تحميل هذا الملف من موقع المناهج الإماراتية



الملف أوراق عمل مراجعة للوحدة التاسعة الغازات

[موقع المناهج](#) ⇨ [المناهج الإماراتية](#) ⇨ [الصف العاشر المتقدم](#) ⇨ [كيمياء](#) ⇨ [الفصل الثالث](#)

روابط مواقع التواصل الاجتماعي بحسب الصف العاشر المتقدم



روابط مواد الصف العاشر المتقدم على تلغرام

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المزيد من الملفات بحسب الصف العاشر المتقدم والمادة كيمياء في الفصل الثالث

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G10 Advanced Chemistry
Chapter 9 - Revision Problems

Multiple Choice Questions.

Q1.	Which of the following equations is Boyles Law?
a.	$P_1/T_1 = P_2/T_2$
b.	$V_1/T_1 = V_2/T_2$
c.	$P_1V_1 = P_2V_2$
d.	$PV = nRT$

Q2.	Which of the following equations is Charles Law?
a.	$P_1/T_1 = P_2/T_2$
b.	$V_1/T_1 = V_2/T_2$
c.	$P_1V_1 = P_2V_2$
d.	$PV = nRT$

Q3.	Which of the following equations is Gay-Lussac Law?
a.	$P_1/T_1 = P_2/T_2$
b.	$V_1/T_1 = V_2/T_2$
c.	$P_1V_1 = P_2V_2$
d.	$PV = nRT$

Q4.	Which of the following equations is the ideal gas law?
a.	$P_1/T_1 = P_2/T_2$
b.	$V_1/T_1 = V_2/T_2$
c.	$P_1V_1 = P_2V_2$
d.	$PV = nRT$

Q5.	What volume is occupied by 1 mol of a gas at STP?
a.	22.4L
b.	1L
c.	44.8L
d.	0L

Q6.	What are STP conditions?
a.	1 atm 298K
b.	1 atm 273K
c.	1 atm 25°C
d.	2 atm 0°C

Q7.	What volume is occupied by 2 mol of a gas at STP?
a.	22.4 L
b.	1 L
c.	44.8 L
d.	0 L

Q8.	What is the volume that 42 g of carbon monoxide gas occupies at STP?
a.	33.6 L
b.	34 L
c.	42 L
d.	0 L

Q9.	Select two conditions under which a gas is least likely to behave ideally.
a.	High pressure
b.	High temperature
c.	Low pressure
d.	Low temperature

Q10.	A 2.00 L flask is filled with ethane gas (C ₂ H ₆) from a small cylinder, with a pressure of 1.08 atm and 15.0 °C. What is the mass of the ethane in the flask?
a.	30.0 g
b.	2.74 g
c.	11.2 g
d.	0.0914 g

Q11.	What is the density of a sample of nitrogen gas (N ₂) that exerts a pressure of 5.30 atm in a 3.50 L container at 125°C?
a.	15.9 g/L
b.	7.95 g
c.	7.95 g/L
d.	15.5 g

Q12.	How many moles of helium gas (He) would be required to fill a 22 L container at a temperature of 35°C and a pressure of 3.1 atm?
a.	2.697 moles
b.	2.70 moles
c.	2.7 moles
d.	0.98 moles

Q13.	What is the volume ratio of carbon monoxide to carbon dioxide in the following equation? $2\text{CO}_{(g)} + 2\text{NO}_{(g)} \rightarrow \text{N}_{2(g)} + 2\text{CO}_{2(g)}$
a.	2:2
b.	1:0
c.	2:1
d.	1:2


Q14.	When 3.00 L of propane (C ₃ H ₈) gas is completely combusted to form water vapor and carbon dioxide at 350°C and 0.990 atm, what mass of water vapor results?
a.	0.232 g
b.	1.05 g
c.	4.18 g
d.	18 g

Q15.	When heated, solid potassium chlorate (KClO ₃) decomposes to form solid potassium chloride and oxygen gas. If 20.8 g of potassium chlorate decomposes, how many litres of oxygen gas will form at STP?
a.	3.80 L
b.	5.7 L
c.	3.8L
d.	5.70 L

Q16a.	Gaseous methane (CH ₄) undergoes complete combustion by reacting with oxygen gas to form carbon dioxide and water vapor. Write a balanced equation for this reaction.
a.	$\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$
b.	$2 \text{CH}_{4(g)} + 4 \text{O}_{2(g)} \rightarrow 2 \text{CO}_{2(g)} + 4 \text{H}_2\text{O}_{(g)}$
c.	$\text{CH}_{4(g)} + 2 \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + 2 \text{H}_2\text{O}_{(g)}$
d.	$\text{CH}_{4(g)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$

Q16b.	Gaseous methane (CH ₄) undergoes complete combustion by reacting with oxygen gas to form carbon dioxide and water vapor. What is the volume ratio of methane to water in this reaction
a.	1:1
b.	1:2
c.	2:1
d.	2:2

Constructed Response Questions.

Q1	<p>Oxygen Consumption: If 5.00 L of hydrogen gas, measured at a temperature of 20.0°C and a pressure of 80.1 kPa, is burned in excess oxygen to form water, what mass of oxygen will be consumed? Assume temperature and pressure remain constant.</p> 
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Q2	<p>A fixed amount of oxygen gas is held in a 1.00 L tank at a pressure of 3.50 atm. The tank is connected to an empty 2.00 L tank by a tube with a valve. After this valve has been opened and the oxygen is allowed to flow freely between the two tanks at a constant temperature, what is the final pressure in the system?</p>
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