مقررات امتحان الإعادة (تمكين ـ للصفوف من 4 - 11) العام الدراسي 2019/2018

لمدارس التعليم العام / الخاص المطبق لمنهاج وزارة التربية والتعليم

مادة الكيمياء	الصف	الفصل الدراسي	عنوان الوحدة	الأقسام	النسبة المئوية للوحدة
	العاشر المسار المتقدم	الثاني	الغازات	قوانين الغازات	
				قانون الغاز المثالي	
				الحسابات الكيميائية للغازات	
		الثالث	المخاليط والمحاليل	أنواع المخاليط	
				تركيز المحاليل	
				العوامل المؤثرة في الذوبان	
			سرعة التفاعلات الكيميائية	نموذج لسرعة التفاعلات	
				العوامل المؤثرة في سرعة التفاعلات	
				قوانين سرعة التفاعلات	
			الاتزان الكيميائي	حالة الاتزان الديناميكي	
				العوامل المؤثرة في الاتزان الكيميائي	
				استخدام ثوابت الاتزان	

مادة الكيمياء	الصف	الفصل الدراسي	عنوان الوحدة	الأقسام	النسبة المئوية للوحدة
	الحادي عشر المسار العام	الثاني	التفاعلات لكيميائية	التفاعلات والمعادلات	
				تصنيف التفاعلات الكيميائية	
				التفاعلات في المحاليل المائية	
		الثالث	المخاليط والمحاليل	أنواع المخاليط	
				تركيز المحاليل	
				العوامل المؤثرة في الذوبان	
			الأحماض والقواعد	مقدمة في الأحماض والقواعد	
				قوة الأحماض والقواعد	
				أيونات الهيدروجين وال PH	

Unit 6 **Quantum Theory and the Electronic Structure of Atoms**	write the electron configuration, orbital notation, and Noble gas notation of elements and ions	
Term 3		
Unit No.	**Title**	**Weight**
Unit 8 **Chemical Bonding I: Basic Concepts**	**Write** Lewis structures for molecules and ionic compounds.	
	Use formal charge to choose the correct Lewis structure for a compound.	
	Write resonance structures for molecules that cannot be represented accurately by only one Lewis structure.	
	Explain how exceptions to the octet rule are caused by an incomplete octet, an odd number of electrons, or more than eight electrons around the central atom.	
	Solve problems involving bond enthalpy.	
Unit 9 **Chemical Bonding II: Molecular Geometry and Hybridization of Atomic Orbitals**	- Represent and identify the molecular geometry of molecules with none, one or more lone pairs in the central atom - Predict and represent the hybridization of an atomic orbital	
	Understand the rules of filling molecular orbitals to maintain or achieve stability	
	Describe bonding in terms of resonance structures and delocalized molecular orbitals	
Unit 10:	- Describe and identify the four types of intermolecular	

Intermolecular Forces and Liquids and Solids	forces - Differentiate between strong and weak interactions of molecules	
	Understand the properties of liquids determined by strength and types of intermolecular forces present	

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الإمارات العربية المتحدة
وزارة التربية والتعليم

Grade: G9 **Stream: ASP** **Subject: Chemistry**

Term 2		
Unit No.	**SLO**	**Weight**
Unit 6 **Ionic compounds and Metals**	understand the properties of ionic compounds	
	describe the properties of metals	
Unit 7 **Covalent Bonding**	- understand electronegativity and its relationship to bond character, be able to interpret results from graphical representations of the relationship - calculate electronegativity and interpret the result - explain how to apply the VSEPR model to predict the molecular shape, bond angle, and hybridization of a molecule	
Unit 8 **Chemical Reactions**	- understand the activity series and how to use it to predict products - predict whether reactions in aqueous solutions will produce a precipitate, water, or a gas	

Term 3		
Unit No	**SLO**	**Weight**
Unit 11 **States of Matter**	- **Describe** viscosity and how it is affected by attractive forces, particle size and shape, and temperature. - **Define** surface tension.	
	- **Describe** the 6 types of phase change and how it involves the addition or removal of energy - **Interpret** phase diagrams	
	Describe the properties of heterogeneous and homogeneous	

& Solutions	mixtures and **provide** examples of each type.	
	Calculate concentration in terms of % by mass, % by volume, molarity, molality, mole fraction	
	Describe the solvation process for ionic and molecular compounds.	

Grade: G10 **Stream: ASP** **Subject: Chemistry**

Term 2		
Unit No.	**Title**	**Weight**
Unit 4 **Gases**	understand, explain, and relate the gas laws proposed by Boyle, Charles, Gay-Lussac, and Avogadro	
	understand and relate the variables of the ideal gas equation	
	- understand, relate, and apply the laws of partial pressures in gases qualitatively and quantitatively - calculate the mass of a gas in an exp*eriment*	
Unit 5 **Thermochemistry**	- understand enthalpy and enthalpy of reactions - apply mathematical routines given thermochemical equations and experimental data and justify results	
	select and apply mathematical routines to apply to calorimetry contexts and justify results	
	apply mathematical routines to determine the standard enthalpy of formation and reaction using both the direct and indirect methods	