

تم تحميل هذا الملف من موقع المناهج المصرية



حل أسئلة الوزارة الأسبوع الثاني

موقع المناهج ← المناهج المصرية ← الصف الأول الإعدادي ← كيمياء ← الفصل الثاني ← حلول ← الملف

تاريخ إضافة الملف على موقع المناهج: 23:18:50 2025-02-21

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المزيد من مادة
كيمياء:

التواصل الاجتماعي بحسب الصف الأول الإعدادي



صفحة المناهج
المصرية على
فيسبوك

المزيد من الملفات بحسب الصف الأول الإعدادي والمادة كيمياء في الفصل الثاني

حل أسئلة الوزارة الأسبوع الأول

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The second week
First Prep
Lesson Two
Unit One
Class Performance

Study the following figures and then answer the following:

1-The two opposite figures show two compounds, one alkaline and the other acidic

A. Which is acidic and which is alkaline?

1 → acidic. 2 → alkaline



B. Which of them results from dissolution of a non-metal and which results from the dissolution of a metal?

1 → non metal. 2 → metal

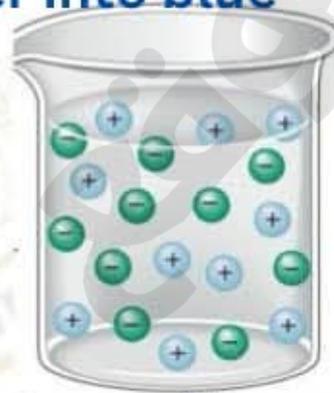
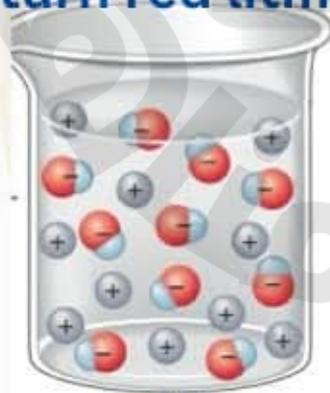
C. What is the result of adding the first tube to the second tube?

Neutralization reaction (forms salt and water)

D. How do you differentiate between them with litmus paper?

1 → turn blue litmus paper into red

2 → turn red litmus paper into blue





The following two figures illustrate two materials:

1-Which is acidic and which is alkaline? And why?

2 → acidic bec it turns blue litmus paper into red

1 → alkaline bec it turns red litmus paper into blue

2- If the acidic substance contains $(2H^+)$,

What is the possible molecular formula of the acid?

H_2CO_3 Carbonic acid
 H_2SO_4 Sulphuric acid
 H_2S Hydrosulphuric acid

3-What are the ions produced from the dissolution of acid in water?

hydrogen cation H^+ and anions of non metal (except O) or negative group (except hydroxide)

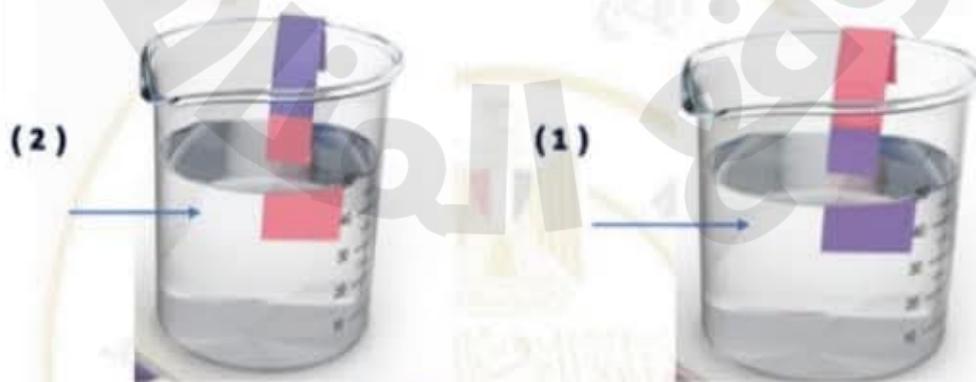
4-If the alkaline substance contains one positive atomic group,

What is the possible molecular formula of the base?

NH_4OH Ammonium hydroxide

5-What are the ions produced from the dissolution of alkali in water?

Cation of metal or positive group and hydroxide anion





The second week

First Prep

Lesson Two

Unit One

Homework

Complete the following:

- 1-The molecular formula of the carbonate group is....., while the molecular formula of the sulphate group is.....
 $(CO_3)^{2-}$
 $(SO_4)^{2-}$
Oxygen
- 2-Oxygenated acids are those that contain element, such as **Carbonic acid** and **Sulphuric acid**
Oxygen **Hydrochloric acid**
- 3-Non-oxygenated acids don't contain element, such as and **hydrobromic acid**
- 4-The number of hydrogen atoms in the acid molecule equals the charge of the **Anion** that forms it, while the number of hydroxide groups in the base molecule equals the charge of the **Cation** that forms it
- 5-Lactic acid provides the muscles with..... during lacking of oxygen, and its accumulation in the muscles causes **Necessary Energy** **muscle cramps**

Give one example of each of the following:

- 1-Positive atomic group
Ammonium group $(NH_4)^+$
- 2-Non-oxygenated acid that forms an anion in a liquid state
Hydrogen bromide (hydrobromic acid) HBr
- 3-Acid secreted by the stomach
Hydrochloric acid HCl
- 4-Oxygenated acid carries three negative charges.
Phosphoric acid H_3PO_4
- 5-Acid secreted by muscles
lactic acid





Write the chemical formula for the following compounds:

- 1-Hydrosulphuric acid H_2S
- 2-Hydrobromic acid HBr
- 3-Sulphurous acid H_2SO_3
- 4-Phosphoric acid H_3PO_4
- 5-Ammonium hydroxide NH_4OH



Write the names of the following chemical compounds and state their type:

- 1- H_2SO_4 Sulphuric acid (oxy acid)
- 2- H_2CO_3 Carbonic acid (oxy acid)
- 3- HCl Hydrochloric acid or hydrogen chloride (non oxy acid)
- 4- HNO_2 Nitrous acid (oxy acid)
- 5- MgO Magnesium oxide (metal oxide)
- 6- SO_3 Sulphite group (polyatomic ion)
- 7- N_2O Nitrous oxide (non metal oxide)

What is the benefit of each of the following:

- 1-Stomach acid Help in digestion of food in stomach
- 2-Lactic acid

It provides the muscles with the necessary energy during their lack of oxygen



The second week
First Prep
Lesson Two
Unit One
Weekly test

What is meant by each of the following:

- 1-Atomic group An ion composed of more than one atom of more than one element acting as one unit carrying a number of + or - charge
- 2-Acids They are substance which dissolves in water giving H^+
- 3-Alkalis They are substance which dissolves in water giving OH^-

How to differentiate between:

- 1- The bicarbonate group and the sulphate group in terms of the number of charges they carry - the molecular formula.

Bicarbonate \rightarrow one negative charge (HCO_3^-)

Sulphate \rightarrow two negative charge (CO_3^{2-})

- 2- Sulphuric acid and calcium hydroxide in terms of molecular formula

Sulphuric acid H_2SO_4
 $Ca(OH)_2$ calcium hydroxide

- 3-Sulphurous acid and sulphuric acid in terms of molecular formula

Sulphurous acid $\rightarrow H_2SO_3$
sulphuric acid $\rightarrow H_2SO_4$

Complete the following sentences:

Arrhenius

- 1-The scientist..... clarified that acids are substances that dissolve in water and produce positive **Hydrogen** ions, while bases, when dissolved in water produce negative **Hydroxide** ions

Hydrogen

- 2-The molecular formula of acids begins with the positive..... cation, and its name is associated with the **Anion** element in its composition



Hydroxide

3-The molecular formula of bases ends with the anion....., and its name is associated with the cation **Positive ion** in its composition

4-The molecular formula of bases that contain a cation Ca^{+2} **$Ca(OH)_2$**

5-The total charge of molecules of any compound equals..... **Zero**

6-The stomach secretes..... **HCl** acid that helps in..... **Digestion**

Write the chemical formula for the following compounds:

1-Hydrochloric acid **HCl**

2-Sulphuric acid **H_2SO_4**

3-Nitric acid **HNO_3**

4-Potassium hydroxide **KOH**

5-Atomic group consisting of three elements **$(HCO_3)^-$**

6-Nitrous acid **HNO_2**



Write the names of the following chemical compounds and state their type:

1- H_2CO_3 **Carbonic acid(oxyacids)**

2- $Mg(OH)_2$ **Magnesium hydroxide (alkali /Base)**

3- H_2S **Hydrogen sulphide /Hydrosulphuric acid
(non oxy acid)**

4- H_3PO_4 **Phosphoric acid
(oxy acid)**