

تم تحميل هذا الملف من موقع المناهج المصرية



حل أسئلة الوزارة الأسبوع الأول

موقع المناهج ← المناهج المصرية ← الصف الثالث الإعدادي ← كيمياء ← الفصل الثاني ← حلول ← الملف

تاريخ إضافة الملف على موقع المناهج: 2025-02-21 23:10:46

ملفات اكتب للمعلم اكتب للطالب | اختبارات الكترونية | اختبارات | حلول | عروض بوربوينت | أوراق عمل
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المزيد من مادة
كيمياء:

التواصل الاجتماعي بحسب الصف الثالث الإعدادي



صفحة المناهج
المصرية على
فيسبوك

المزيد من الملفات بحسب الصف الثالث الإعدادي والمادة كيمياء في الفصل الثاني

مراجعة مهمة جداً للوحدة الأولى

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Classroom performance

Lesson 1: Chemical reaction , Thermal Decomposition Reactions

Question 1: Write the Scientific Term:

- 1) The breaking bonds between the molecules of the reactants and forming new bonds between the molecules of the molecules resulting from the reaction. **Chemical reactions**
- 2) Chemical reactions in which the thermal compound decomposes into its primary elements or simple compound. **Thermal decomposition**
- 3) A safety means is found in modern cars and has a form of an inflatable bag. **Air bags**

Question 2: Choose the correct answer:

- 1) When heating copper hydroxide, we get
(copper carbonate and water vapor - copper oxide and water vapor - copper and hydrogen - copper oxide and hydrogen)
- 2) Some metal nitrates decompose with heat into
(metal nitrite and oxygen - metal nitrate and oxygen - nitric oxide and hydrogen - metal nitrate)
- 3) When heating red mercury oxide, it decomposes into
(oxygen - mercury - oxygen and mercury - there is no correct answer)
- 4) Almost metal carbonates decomposed by heating into metals and
(O2 - NO2 - CO2 - SO3)

Help in burning gasoline and get power
production of medicines & fertilizers

Question 3: What is the importance of the chemical reactions?

1-thermal decomposition

Question 4: What are the types of chemical reactions?

2-substitution

Question 5: according to your study

3-oxidation & reduction

Mention the names of the chemical compounds that give oxygen gas when thermally decomposed.

**Mercuric oxide HgO
sodium nitrate NaNO3**



Homework Performances



Lesson 1: Chemical reactions

Thermal decomposition reactions

Question 1: Complete the following statements:

1) The compound decomposes with heat into its simple components in reactions of

Thermal decomposition

Sulphate

Metal oxide

2) Most metals decomposed by heating into and sulfur trioxide gas

3) **Airbags** is considered one of the most important safety means in modern cars, as it is filled with gas **nitrogen**

4) **CO₂** gas turbid the clear lime water, while **O₂** gas increases the glow of glowing splint.

Question 2: Give reasons for the following:

1) A silver color appears when red mercury oxide is heated.

red mercuric oxide is decomposed by heat into silver mercury and oxygen gas

2) A black substance is formed when green copper carbonate is heated strongly.

green copper carbonate decompose by heat into black copper oxide and CO₂ evolves

3) A black color appears when blue copper sulfate is heated.

blue copper sulphate decompose by heat into black copper oxide and SO₃ evolves

4) The color of sodium nitrate changes to yellowish white by heating.

sodium nitrate is decomposed by heat into yellowish white sodium nitrite and oxygen

Question 3: What happens in the following cases:

1) Heating red mercury oxide.

red mercuric oxide is decomposed by heat into silver mercury and O₂ evolves

2) Heating blue copper hydroxide.

blue copper hydroxide decompose by heat into black copper oxide and water vapor

3) Heating green copper carbonate.

green copper carbonate decompose by heat into black copper oxide and CO₂ evolves

4) Heating white sodium nitrate

sodium nitrate is decomposed by heat into yellowish white sodium nitrite and oxygen



weekly Performances

Lesson 1: Chemical Reactions Thermal decomposition reactions

Question 1: Choose the correct answer:

1) One of the benefits of chemical reactions in our lives is

- A- manufacturing of medicine B- artificial fibers industry
C- Fertilizer industry D- All of the above

2) When blue copper hydroxide is heated, it decomposes into

- A- Copper oxide and hydrogen B- Copper oxide and water vapor
C- Copper and oxygen D- Hydrogen and oxygen

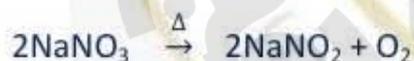
3) Carbon dioxide gas is evolved when the is decomposed by heating.

- A -Cu(OH)₂ B- CuCO₃ C- CuSO₄ D- HgO

4) Thermal decomposition of blue copper sulfate gives copper oxide and gas

- A- Sulfur dioxide B- Sulfur trioxide C- Oxygen D- Sulfur

5) The following chemical equation :-



represents a reaction of

- A- Simple substitution B- Double substitution C- Thermal decomposition
d- Neutralization

6) The airbag contains the sodium compound

- A- Sulfate B- Azide C- Oxide D- Carbonate

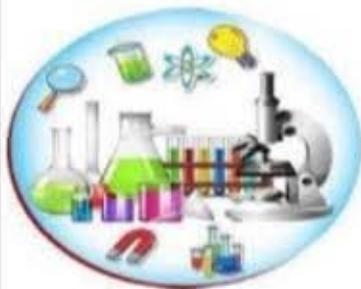
Question 3: What happens when?

electric spark is caused which decompose sodium azide into sodium and hydrogen gas

1) A rapid and sudden decrease in the speed of the car that fills the air bag

2) Heating an amount of white sodium nitrate

sodium nitrate is decomposed by heat into yellowish white sodium nitrite and oxygen



It increases the glowing

3) A glowing splint is exposed to the gas resulting from the thermal decomposition of HgO

Because green copper carbonate decomposes by heat into black copper oxide and CO_2 evolves

Question 3: Give reasons for the following:

1) A black substance is formed when green copper carbonate is heated strongly

2) A yellowish-white substance is formed when white sodium nitrate is heated

Because sodium nitrate is decomposed by heat into yellowish white sodium nitrite and oxygen

3) When heating mercury oxide, its mass decreases and its color changes

Because red mercuric oxide is decomposed by heat into silver mercury and oxygen gas

4) Copper hydroxide turns black when heated

Because blue copper hydroxide decomposes by heat into black copper oxide and water vapor

5) Chemical reactions are of great importance in our daily lives

6) The airbag is considered one of the

Help in burning gasoline and get power
production of medicines & fertilizers

Because they get inflated on the occurrence of the car crash due to decomposition of sodium azide NaN_3 into sodium and nitrogen gas evolves filling the air bag